Observations:

Initial unknown chloride sample weight (g):
Weight of AgCl precipitate:
Approximate volume of AgNO$_3$ solution added (ml):

Moles of AgCl = (AgCl precipitate mass/143.32):
Weight of Cl$^-$ in XCl sample = (moles of AgCl)x(atomic weight of Cl$^-$ = 35.4527) :
% Cl$^-$ in sample = (weight of Cl$^-$ in XCl sample / weight of XCl sample) x 100 % :

What happened as excess AgNO$_3$ solution was added to the XCl solution ?